Fuel Cells and the Emerging Hydrogen Energy Economy

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Presented to

Association of Energy Engineers

Southern California Chapter 2003 Annual Expo and Conference

20 March 2003



Overview

- The Promise of Fuel Cells
 - What fuel cells are and are not
- Types of Fuel Cells
- Fuel Cell Applications
- Availability of Fuel Cell Power Systems
- Economics of Fuel Cell Power Systems
- Summary



Fuel cells are not:
... a panacea for the future
... more efficient than all other technologies
... free of pollutants and greenhouse gases

... the solution to all our energy and environmental concerns



- Fuel cells are:
 - ... An important part of a balanced future energy policy
 - ... A viable alternative to batteries, ICEs and other energy technologies in many applications
 - ... More efficient than conventional energy technologies in specific applications
 - ... A technology whose time has come



- Polymer Electrolyte Membrane (PEM)
- Direct Methanol
- Alkali
- Metal-air
- Phosphoric Acid
- Molten Carbonate
- Solid Oxide



- Handheld applications
- Portable power
- Small-scale distributed power
- Large-scale distributed power
- Transportation, mobility & utility

- Metal-air fuel cells
 Direct methanol fuel cells
 - Cell phones
 - Lap top computers
 - PDAs, calculators
 - Diagnostic equipment
 - Power tools
 - President George W. Bush speaking on a cell phone powered by an *MTI MicroFuel Cells* direct methanol fuel cell



- Polymer electrolyte membrane
- Metal-Air fuel cells
- Alkali fuel cells
 - Remote applications
 - Recreational use
 - Emergency response
 - Backup power



Metallic Power 1.6 kW zinc-air fuel cell

Fuel Cell Applications Small-scale distributed power

- Polymer electrolyte membrane
- Metal-air fuel cells
- Solid oxide fuel cells

- Residential
- Small Commercial
- Back up or primary power
- Grid parallel
- Grid independent
- Pipeline NG or LPG operation SOFC stack



Global Thermoelectric 800W

- Solid oxide
- Molten carbonate
- Phosphoric acid
 - Large commercial & industrial applications
 - Back up or primary power
 - Premium power
 - Natural gas, LPG or digester gas
 - Biomass under development

Fuel Cell Applications Transportation & mobility

- Polymer electrolyte membrane
- Alkali fuel cells
- Metal air fuel cells
- Solid oxide fuel cells
 - Automobiles
 - Public transportation
 - Hotel power (trucks and trains)
 - Utility vehicles
 - Personal mobility vehicles

Fuel Cell Availability Commercially available fuel cells

- PEM fuel cells
 - Small distributed power/cogen
 - Portable power
- Phosphoric acid fuel cells
 - Commercial/industrial DPG/cogen/renewable
- Molten carbonate fuel cells
 - 250 kW– 2 MW scale DPG/cogen/combined cycle
 - Renewables on the horizon



Fuel Cell Availability

- Plug Power
 - 4.5 kWe back up power/cogen
 - NG operation
- Ballard/Coleman Powermate
 - 1 kW backup/portable system
 - Hydrogen operation
- Avista Laboratories
 - 100, 500 W, 1 kW
 - Hydrogen operation



Fuel Cell Availability EM fuel cells

Plug Power GenSys Residential/light commercial



- 4.5 kWe/6.7 kWth
- Backup power
- Grid parallel or independent
- NG (hydrogen)
- 1- and 2-year warrantees



- Coleman Powermate
 - Ballard PEM fuel cell
 - -1 kWe
 - -120 VAC
 - Transportable
 - Requires hydrogen
 - Industrial only





Fuel Cell Availability

- Avista Labs
 - Independence 1000
 - Commercial/industrial only
 - 48 VDC output
 - Premium power



Fuel Cell Availability Phosphoric acid fuel cells

• UTC Fuel Cells

- 200 kW industrial operation
- 900,000 Btu/hr cogen heat
- Shown in digester gas application



Fuel Cell Availability Molten carbonate fuel cell

Fuel Cell Energy

- 250 kW, 1 MW, 2 MW units available
- 47 50 % efficient
- Cogen heat available at > 400 °C
- Renewable fuel systems under development





Fuel Cell Economics

- PEM Fuel Cells
 - \$6,000/kW hydrogen only
 - \$11,000/kW NG (fuel processor) with cogen
 - \$13,800/kW installed
 - Expected to drop to \$4,500/kW in 2004
- PAFC
 - \$4,000/kW NG with cogen
 - \$5,200/kW installed
- MCFC
 - \$4,000/kW NG with cogen
 - \$4,700/kW installed



California

- Various state and local programs offering up to \$4,500/kW (50 %) for renewable fuel sources, \$2,500/kW (40 %) for NG and LPG operation
- DoD Climate Change Fuel Cell Program – \$1,000/kW (33 %)
- Federal Business Tax Credit (S.461)
 - -\$1,000/kW



Fuel Cell Economics Example case - PAFC



Fuel Cell Economics Example case - PAFC





Summary

- Fuel cell technology is available today
- Economics are very site specific
 - Rely heavily on rebates, tax credits
- Value beyond simple power more difficult to quantify
 - Cost of outages
 - Value of "security of supply"



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Addendum

Introduction to Fuel Cells

- Fuel Cells 101
 - Science & Technology of Fuel Cells
- Fuel cells 102
 - Types of Fuel Cells
 - Major Distinctions

Fuel Cells 101 Science and technology of fuel cells

- History of Fuel Cells
 - Invented in 1839 by Sir William Grove
 - Known then as the gas voltaic battery
 - The term *fuel cell* was proposed by Mond and Langer in 1889
 - Later revived for space programs
 - Evolved into seven major commercially viable technologies
 - Several other technologies waiting in the wings

Fuel Cells 101 Science and technology of fuel cells

- Fuel Cell Science
 - Fuel and oxidant combined to produce electric current
 - Fuel delivered at the anode
 - Oxidant delivered at the cathode
 - Hydrogen-oxygen fuel cell has the simplest chemical system

Anode reaction: $H_2 \rightarrow 2 H^+ + 2 e^-$ Cathode reactions: $\frac{1}{2} O_2 + 2 e^- \rightarrow O^{2-}$
 $O^{2-} + 2 H^+ \rightarrow H_2O$

Fuel Cells 101 Science and technology of fuel cells

H_-

 H_2

 H_2^-

Catalyst

 $H_2 \rightarrow$

 H_2

 H_{2}

 H_{2}

Fuel Cell Science

- 1) Hydrogen reacts at the anode to produce protons and free electrons
- 2) Electrons are conducted across the "load" between the anode and cathode
- 3) Protons are conducted across the electrolyte
- 4) Oxygen molecules react with electrons to form oxide anions on the cathode
 Anode
- 5) Protons and oxide anions react to form water vapor

Membrane (electrolyte)

O²⁻

 O^{2}

02-

02-

H₂C

H₂(

Cathode Catalyst

H₂O

 H_2O

- Types of fuel cells
 - Ways of classifying fuel cells
- Major distinctions
 - Temperature
 - Electrolyte
 - Fuel type

- Ways of Classifying Fuel Cells
 - Temperature
 - Low, intermediate, high
 - Electrolyte
 - Solid, liquid
 - Acidic, basic
 - Fuel
 - Hydrogen
 - Syngas (internally reforming)
 - Metallic

Fuel Cells 102 Types of fuel cells and their major distinctions

- Low temperature (< 100^{-o}C)
 - PEM
 - Direct Methanol
 - Alkali
 - Metallic
- Intermediate temperature (100 250 °C)
 - Phosphoric acid
 - Alkali
- High temperature (up to 1000 °C)
 - Molten carbonate
 - Solid oxide

- Solid electrolyte
 - Solid oxide
 - PEM
 - Direct Methanol
- Liquid electrolyte
 - Metallic
 - Phosphoric acid
 - Molten carbonate
 - Alkali

- Acid electrolyte
 - PEM, Direct Methanol, Phosphoric Acid
 - Proton is charge carrier
- Basic electrolyte
 - Alkali, Molten carbonate, Metallic, Solid oxide
 - Anionic charge carrier
 - OH⁻, CO₃²⁻, OH ⁻, O²⁻, respectively

- Fuel Type
 - Metallic
 - Aluminum, zinc
 - Hydrogen
 - PEM, Alkali, Phosphoric acid
 - Syngas
 - Molten carbonate, Solid oxide
 - Alcohol
 - Direct methanol